MODULE HANDBOOK

Module Name	Main Elements of Inorganic Chemistry
Module Level	Bachelor
Abbreviation, if applicable	3074213046
Sub-heading, if applicable	-
Course included in the module, if applicable	-
Semester/term	6^{th}/ third year
Module coordinator(s)	Dina Kartika Maharani, S.Si., M.Sc
Lecturer(s)	1. Dr. Achmad Lutfi, M.Pd. 2. Dr. Muchlis, M.Pd. 3. Dina Kartika M., S.Si., M.Sc, 4. Kusumawati Dwiningsih, S.Pd., M.Pd. 5. Rusly Hidayah, S.Si., M.Pd.
Language	Bahasa Indonesia
Classification within the curriculum	Compulsory
Teaching format/class hours per week during the semester	3 hours lectures (50 min / hour)
Workload	1 CU for bachelor degree equals to 3 workhours per week or 170 minutes (50' face to face learning, 60' structured learning, and 60' independent learning). In one semester, courses are conducted in 14 weeks (excluding mid and end-term exam). Thus, 1 CU equals to 39.67 workhours per semester. One CU equals to 1.59 ECTS.
Credit point	3 CU = 3 x 1.59 = 4.77 ECTS
Requirement	General Chemistry II
Learning Outcomes	**General Competence (knowledge)**: Students can mastering theoretical concepts on the structure, dynamics and energy of chemicals, as well as the basic principles of separation, analysis, synthesis and characterization of main group elements **Specific Competence** : At the end of the lecture, students can understand the position, physico-chemical properties, laboratory manufacture, types of compounds and their uses of Alkali Metals, Alkaline Earth metals, Boron, Carbon, Nitrogen, Oxygen, Halogens, Noble Gases.

Content	Course materials discuss the understanding of Role and status of theory in Inorganic chemistry, Origin of elements, Classification of elements in the periodic system; Hydrogen and its compounds: Position in the periodic table, Physical and chemical properties, Isotopes of hydrogen, Hybrids of elements, Water and related matters; Source and extraction, physico-chemical properties and uses, manufacture, properties and uses: Alkali Metals, Alkaline Earth metals, Boron, Carbon, Nitrogen, Oxygen, Halogens, Noble Gases
Study/exam achievements	Students are considered to be competent and pass if at least get 55 Final score is calculated as follows: 20% assignment + 30% Task + 20% Table index of graduation • A = 4 (85 ≤-≥ 100) • A- = 3,75 (80 ≤-< 85) • B+ = 3,5 (75 ≤- < 80) • B = 3 (70 ≤-< 75) • B- = 2,75 (65 ≤-<75) • C+ = 2,5 (60 ≤-<65) • C = 2 (55 ≤-<60) • D = 1 (40 ≤-<55) • E = 0 (0 ≤-<40)
Forms of media	Computer, LCD, White board
Learning Methods	Lectures, discussion, assignment
Literature	1. Lee, J.D. 1991. *Concise Inorganic Chemistry*. Four Edition. London: Chapman & Hall. 2. Madan, R.D. 1997. *Modern Inorganic Chemistry*. New Delhi: S. Chand and Company LDT. 3. Miesler, Fischer, Tarr. 2014. Inorganic Chemistry. 5th . Pearson 4. Education Inc.Perry, Dale L. 2011. *Handbook of Inorganic Compounds, Second Edition* **(**Hardcover) – May 18, 2011. ISBN-13: 000-1439814619 ISBN-10: 14398146