MODULE HANDBOOK

Module Name	Practicum of Inorganic Chemistry
Module Level	Bachelor
Abbreviation, if applicable	3074211048
Sub-heading, if applicable	-
Course included in the module, if applicable	-
Semester/term	6th/ Third Year
Module coordinator(s)	Kusumawati Dwiningsih, S.Pd., M.Pd.
Lecturer(s)	**Dr. Achmad Lutfi, M.Pd.; Dr. Amaria, M.Si., Prof. dr. Sari Edi C., M.Si, Dr. Muchlis, M.Pd.; Dina Kartika M., S.Si., M.Sc, Kusumawati D., S.Pd.,M.Pd.; Rusly Hidayah, S.Si., M.Pd.**
Language	Bahasa Indonesia
Classification within the curriculum	Compulsory Course
Teaching format/class hours per week during the semester	2 hours lectures (150 min / hour)
Workload	2 x 50 minutes lectures, 2 x 60 minutes structured activity, 2 x 60 minutes individual activity, 14 weeks per semester, 79,33 total hours per semester ~ 3.18 ECTS**
Credit point	2 CU x 1.59 = 3.18 ECTS
Requirement	General Chemistry II
Learning Outcomes	**General Competence (knowledge)**: Students can understand the physical-chemical properties, preparation of main group elements (alkalis, alkaline earth, boron family, carbon family, nitrogen family, oxygen family, halogen and hydrogen) and transition elements **Specific Competence** : At the end of the lecture, students can study physical-chemical properties, preparation of its compounds in laboratory scale of alkalis, alkaline earth, boron family, carbon family, nitrogen family, oxygen family, halogen and hydrogen, study preparation of cis trans metal complexes, metal salt complexes, study the strength of ligand fields in metal complexes, and study reactions in metal complexes
Content	Course materials discuss physical-chemical properties, preparation of its compounds in laboratory scale of alkalis, alkaline earth, boron family, carbon family, nitrogen family, oxygen family, halogen and hydrogen, study preparation of cis trans metal complexes, metal salt complexes, study the strength of ligand fields in metal complexes, and study reactions in metal complexes.

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<tr><td>Study/exam achievements</td><td>Students are considered to complete the course and pass if they obtain at least 40% of maximum final grade. The final grade (NA) is calculated based on the following ratio:
<table>
<tr><th>Assessment Components</th><th>Percentage of contribution</th></tr>
<tr><td>Participation</td><td>20%</td></tr>
<tr><td>Assignment</td><td>30%</td></tr>
<tr><td>Mid-semester test</td><td>20%</td></tr>
<tr><td>Final semester test</td><td>30%</td></tr>
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<tr><td>Media:</td><td>Computer, LCD, White board, Chemical Equipment</td></tr>
<tr><td>Learning Methods</td><td>Individual assignment, group assignment, discussion, and presentation</td></tr>
<tr><td>Literature:</td><td>1. Lee, J.D. 1991. ConciseInorganic Chemistry. Four Edition. London: Chapman & Hall.
2. Madan, R.D. 1997. Modern Inoragnic Chemistry. New Delhi: S. Chand and Company LDT.
3. Manku, G.S. 1980.Inorganic Chemistry. India: Tata Mc Graw Hill Book Co.
4. Sugiarto, B. dkk. 1997. Kimia Anorganik. Surabaya: Unipress IKIP Surabaya</td></tr>
<tr><td rowspan="2">Notes:</td><td>*1 CU in learning process = three periods consist of: (a) scheduled instruction in a classroom or laboratory (50 minutes); (b) structured activity (60 minutes); and (c) individual activity (60 minutes) according to the Regulation of Indonesia Ministry of Research, Technology, and Higher Education No. 44 Year 2015 jo. The Regulation of Indonesia Ministry of Research, Technology, and Higher Education No. 50 Year 2018.</td></tr>
<tr><td>**1 CU = 1,59 ECTS according to Rector Decree Of Universitas Negeri Surabaya No. 598/Un38/Hk/Ak/2019</td></tr>
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