

## MODULE HANDBOOK

Modul Name	<b>Laboratory Experimental of Inorganic Chemistry</b>
Module Level	Bachelor
Abbreviation, applicable if	3074211051
Sub-heading, applicable if	-
Course included in the module, if applicable	-
Semester/term	6 <sup>th</sup> / third year
Modul coordinator(s)	Dina Kartika Maharani, S.Si., M.Sc
Lecturer(s)	Dr. Achmad Lutfi, M.Pd.; Dr. Amaria, M.Si., Prof. dr. Sari Edi C., M.Si, Dr. Muchlis, M.Pd.; Dina Kartika M., S.Si., M.Sc, Kusumawati D., S.Pd.,M.Pd.; Rusly Hidayah, S.Si., M.Pd.
Language	Bahasa Indonesia
Classification within the curriculum	Compulsory
Teaching format/class hours per week during the semester	3 hours lectures (50 min / hour)
Workload	1 CU for bachelor degree equals to 3 workhours per week or 170 minutes (50' face to face learning, 60' structured learning, and 60' independent learning). In one semester, courses are conducted in 14 weeks (excluding mid and end-term exam). Thus, 1 CU equals to 39.67 workhours per semester. One CU equals to 1.59 ECTS
Credit point	3 CU = 3 x 1.59 = 4.77 ECTS
Requirement	General Chemistry II
Learning Outcomes	<p><b>General Competence (knowledge):</b> Students can understand the physical-chemical properties, preparation of main group elements (alkalis, alkaline earth, boron family, carbon family, nitrogen family, oxygen family, halogen and hydrogen) and transition elements</p> <p><b>Spesific Competence :</b> At the end of the lecture, students can study physical-chemical properties, preparation of its compounds in laboratory scale of alkalis, alkaline earth, boron family, carbon family, nitrogen family, oxygen family, halogen and hydrogen, study preparation of cis trans metal complexes, metal salt complexes, study the strenght of ligand fields in metal complexes, and study reactions in metal complexes</p>
Content	Course materials discuss physical-chemical properties, preparation of its compounds in laboratory scale of alkalis, alkaline earth, boron family, carbon family, nitrogen family, oxygen family, halogen and hydrogen, study preparation of cis trans metal complexes, metal salt complexes, study the strenght of ligand fields in metal complexes, and study reactions in metal complexes.
Study/exam achievements	Students are considered to be competent and pass if at least get 55  Final score is calculated as follows: 25% experiment + 20% Task + 20%

	<p>middle exam (UTS)&amp; 35% final exam (UAS)</p> <p>Tableindex of graduation</p> <ul style="list-style-type: none"> <li>• 0 - 44,999 E,</li> <li>• 45 - 54,999 D,</li> <li>• 55 - 69,999 C,</li> <li>• 70 - 74,999 B-,</li> <li>• 75 - 79,999 B,</li> <li>• 80 - 84,999 B+,</li> <li>• 85 - 100 A</li> </ul>
Forms of media	Laboratory experiment, white board, LCD
Learning Methods	Lectures, discussion, assignment, experiment
Literatur	<ol style="list-style-type: none"> <li>1. Lee, J.D. 1991. <i>ConciseInorganic Chemistry</i>. Four Edition. London: Chapman &amp; Hall.</li> <li>2. Madan, R.D. 1997. <i>Modern Inoragnic Chemistry</i>. New Delhi: S. Chand and Company LDT.</li> <li>3. Manku, G.S. 1980.<i>Inorganic Chemistry</i>. India: Tata Mc Graw Hill Book Co.</li> <li>4. Miesler, Fischer, Tarr. 2014. <i>Inorganic Chemistry</i>. 5<sup>th</sup> . Pearson Education Inc.</li> </ol>