

# **PORTFOLIO BASIC CHEMISTRY I**

**ACADEMIC YEAR 2019/2020 ODD SEMESTER**



**Course Coordinator:  
Dr. Harun Nasrudin, M.S.**

**Teaching Team:  
Dr. Utiya Azizah, M.Pd.; Rusly Hidayah, S.Si., M.Pd.;  
Prof. Suyatno, M.Si.; Dr. Maria Monica SBW, M.Si.; Dr. Nuniek Herdyastuti,  
M.Si.; Dr. Amaria, M.Si., Prof. Sari Edy C. M.Si.**

**CHEMISTRY DEPARTMENT  
FACULTY OF MATHEMATICS AND SCIENCE  
UNIVERSITAS NEGERI SURABAYA**

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## A. SEMESTER LEARNING ACTIYITY PLAN

### A.1. COURSE IDENTITY

Module Name	Basic Chemistry I
Module level	Bachelor
Abbreviation, if applicable	8420403123
Sub-heading, if applicable	-
Course included in the module, if applicable	-
Semester/term	1 <sup>st</sup> /First Year
Module coordinator(s)	Dr. Harun Nasrudin,M.S.
Lecturer(s)	Dr. Harun Nasrudin,M.S.; Dr. Utiya Azizah, M.Pd.; Rusly Hidayah, S.Si., M.Pd.; Prof. Suyatno, M.Si.; Dr. Maria Monica SBW, M.Si.; Dr. Nuniek Herdyastuti, M.Si.; Dr. Amaria, M.Si., Prof. Sari Edy C. M.Si.
Language	Indonesian
Classification within the curriculum	Compulsory Course
Teaching format/class hours per week during the semester:	3 hours lecturers (50 min per hours)
Workload:	1 CU for bachelor degree equals to 3 workhours per week or 170 minutes (50' face to face learning, 60' structured learning, and 60' independent learning). In one semester, courses are conducted in 14 weeks (excluding mid and end-term exam). Thus, 1 CU equals to 39.67 workhours per semester. One CU equals to 1.587 ECTS.
Credit points:	3 CU (4,761 ECTS)
Prerequisites course(s):	-
Targeted learning outcomes:	<p>CLO 1 Students have the ability to utilize learning resources and ICT to support mastery of concepts and theories of the scientific method, material properties, stoichiometry, atomic structure, system periodic Elements, chemical bonds, energetics, and solutions.</p> <p>CLO 2 Students have the ability to make decisions about the relationship of basic concepts chemistry with laboratory activities and presence chemistry in everyday life.</p> <p>CLO 3 Students have knowledge of the scientific method, material properties, stoichiometry, atomic structure, system periodic elements, chemical bonds, energetics, and solutions.</p> <p>CLO 4 Students have the ability to have an honest and responsible attitude in carry out lectures and practicum.</p>


Content:	<p><b>Introduction:</b> The stages of the scientific method, Chemistry as a scientific activity, material and energy, extensive and intensive properties, chemical and physical properties, elements, compounds, and mixtures</p> <p><b>Stoichiometry:</b> Basic Chemistry Law, Atoms and Molecules, Mole Concepts, Avogadro Constanta, Compound Formulas, Chemical Reactions and Equalization, Polarity and Equivalentents</p> <p><b>Atomic Structure:</b> Basic Particles, Hydrogen Atom Spectrum and Rutherford Atomic Model, Bohr Atomic Model, Atomic Wave Mechanics Model, Electron Configuration</p> <p><b>Periodic System of Elements:</b> Development of the Periodic System, Periodic System and Electron Configuration, Periodicity Properties (Atomic Radius, Ionization Energy, Electron Affinity, and Electronegativity)</p> <p><b>Chemical Bonds:</b> Ion Bonds, Covalent Bonds, Molecular Structures, Metal Bonds, and Chemical Styles (London Style v.d Waals, Hydrogen Bonds,)</p> <p><b>Energetics:</b> Several Terms (Systems, environment, state functions, adiabatic processes, isotherm processes, work, heat capacity, etc.), Law I Thermodynamics, Hess Law, Bonding Energy, Thermochemistry, Law II Thermodynamics, Entropy, Free Energy.</p> <p><b>Solution:</b> Electrolyte and non-electrolyte solution, colligative properties, acid-base, pH of solution, hydrolysis, namesake ion, buffer solution, and titration.</p>
Study / exam achievements:	<p>Students are considered to be competent and pass if at least get 55</p> <p>Final score is calculated as follows: 20% participation + 30% assignment + 20% middle exam (UTS) &amp; 30% final exam (UAS)</p> <p>Table index of graduation</p> <ul style="list-style-type: none"> <li>• A = 4 (85 - 100)</li> <li>• A- = 3,75 (80 - 85)</li> <li>• B+ = 3,5 (75 - 80)</li> <li>• B = 3 (70 - 75)</li> <li>• B- = 2,75 (65 - 75)</li> <li>• C+ = 2,5 (60 - 65)</li> <li>• C = 2 (55 - 60)</li> <li>• D = 1 (40 - 55)</li> <li>• E = 0 (0 - 40)</li> </ul>
Media:	Computer, LCD, White board
Learning Methods	Individuals assignment, group assignment, discussion, presentation, and practicum
Literature:	1. Tim Kimia Dasar. 2017. <i>Kimia Dasar I</i> . Surabaya: Unesa University Press.

	<ol style="list-style-type: none"> <li>2. Brady and Humiston. 2004. <i>General Chemistry, Principles and Structures</i>. New York: John Willey and Sons.</li> <li>3. Chang, Raymond. 2005. <i>General Chemistry The Essential Concepts Third Edition</i>. USA: McGraw Hill.</li> <li>4. Achmad, Hiskia dan Tupamahu. 1990. <i>Penuntun Belajar Struktur Atom, Struktur Molekul, Sistem Periodik</i>. Bandung: ITB.</li> <li>5. Achmad, Hiskia dan Tupamahu. 1991. <i>Stoikiometri dan Energetika Kimia</i>, Bandung, PT Citra Aditya Bakti.</li> <li>6. Ahmad, Hiskia. 1990. <i>Kimia Larutan</i>. Bandung: Jurusan Kimia FMIPA ITB</li> </ol>
Note	Basic chemistry 1 covers the activities of theory, practicum and presentation.

## A.2. COURSE TOPIC

This course examines the concept of Scientific Methods, Material Properties, Stoichiometry, Atomic Structure, Periodic System of Elements, Chemical Bonding, Energetics, Solutions, and appropriate laboratory activities through discussion, assignment, and practicum.

### A.3. COURSE PROGRAM

	<b>UNIVERSITAS NEGERI SURABAYA</b> <b>FACULTY OF MATHEMATICS AND NATURAL SCIENCE</b> <b>UNDERGRADUATE PROGRAMME OF CHEMISTRY EDUCATION</b>					<b>Document Code</b>
	<b>SEMESTER LEARNING ACTIYITY PLAN</b>					
<b>COURSE</b>	<b>CODE</b>	<b>Course Group</b>	<b>Credit Unit</b>		<b>Semester</b>	<b>Date</b>
BASIC CHEMISTRY 1	8420403123		T= 2	P= 1	1	November 30, 2019
<b>AUTHORIZATION CHEMISTRY EDUCATION</b>	<b>Compiler</b>		<b>Coordinator</b>		<b>Head of Study Program</b>	
	Dr. Harun Nasrudin, M.S.		Dr. Nuniek Herdyastuti, M.Si.		Dr. Sukarmin, M.Pd	
<b>Learning Outcomes</b>	<b>Program Learning Outcomes (PLO)</b>					
	PLO1 (KNO-1)	Mastering the theoretical concepts of structure, dynamics, and energy, as well as the basic principles of separation, analysis, synthesis, and characterization of chemicals				
	PLO3 (SKI-1)	Mastering the principles of occupational health and safety, managing laboratories, using the equipment and operating chemical instruments				
	<b>Course Learning Outcomes (CLO)</b>					
	CLO1	Students have the ability to utilize learning resources and ICT to support mastery of concepts and theories of the scientific method, material properties, stoichiometry, atomic structure, system periodic.				
	CLO2	Students have the ability to make decisions about the relationship of basic concepts chemistry with laboratory activities and presence chemistry in everyday life.				
	CLO3	Students have knowledge of the scientific method, material properties, stoichiometry, atomic structure, system periodic elements, chemical bonds, energetics, and solutions.				
	CLO4	Students have the ability to have an honest and responsible attitude in carry out lectures and practicum.				
	<b>Sub CLO</b>					
Sub-CLO1	Describes chemistry as the result of scientific activities that study matter with universal properties					

	Sub-CLO2	Applying the things that underlie stoichiometry, namely: basic laws of chemistry, atoms and molecules, the concept of moles and Avogadro's constant, compound formulas, chemical reactions and polarity and equivalence to complete chemical calculations
	Sub-CLO3	Analyzed the development of the discovery and the elementary particles of the atom according to Rutherford, Bohr, wave mechanics and electron configuration
	Sub-CLO4	Analyze the development, usefulness, and basis for composing the periodic system and its relation to the electron configuration of the elements and their periodic properties
	Sub-CLO5	Identify the relationship between chemical bonds and chemical forces to explain knowledge according to the study program.
	Sub-CLO6	Describing terms, the laws of thermodynamics, and determining the occurrence of thermodynamic reactions
	Sub-CLO7	Analyze several aspects of the solution and apply them in quantitative terms
<b>Brief Description of the Course</b>	Study of basic concepts: Scientific Methods, Material Properties, Stoichiometry, Atomic Structure, Periodic System of Elements, Chemical Bonding, Energetics, Solutions, and appropriate laboratory activities through discussion, assignment, and practicum.	
<b>Study Materials: Learning Materials</b>	<p><b>Introduction:</b> The stages of the scientific method, Chemistry as a scientific activity, material and energy, extensive and intensive properties, chemical and physical properties, elements, compounds, and mixtures</p> <p><b>Stoichiometry:</b> Basic Chemistry Law, Atoms and Molecules, Mole Concepts, Avogadro Constanta, Compound Formulas, Chemical Reactions and Equalization, Polarity and Equivalent</p> <p><b>Atomic Structure:</b> Basic Particles, Hydrogen Atom Spectrum and Rutherford Atomic Model, Bohr Atomic Model, Atomic Wave Mechanics Model, Electron Configuration</p> <p><b>Periodic System of Elements:</b> Development of the Periodic System, Periodic System and Electron Configuration, Periodicity Properties (Atomic Radius, Ionization Energy, Electron Affinity, and Electronegativity)</p> <p><b>Chemical Bonds:</b> Ion Bonds, Covalent Bonds, Molecular Structures, Metal Bonds, and Chemical Styles (London Style v.d Waals, Hydrogen Bonds,)</p> <p><b>Energetics:</b> Several Terms (Systems, environment, state functions, adiabatic processes, isotherm processes, work, heat capacity, etc.), Law I Thermodynamics, Hess Law, Bonding Energy, Thermochemistry, Law II Thermodynamics, Entropy, Free Energy.</p> <p><b>Solution:</b> Electrolyte and non-electrolyte solution, colligative properties, acid-base, pH of solution, hydrolysis, namesake ion, buffer solution, and titration.</p>	
<b>Reference</b>	<b>Main :</b>	
		<ol style="list-style-type: none"> <li>1. Tim Kimia Dasar. 2017. <i>Kimia Dasar I</i>. Surabaya: Unesa University Press.</li> <li>2. Brady and Humiston. 2004. <i>General Chemistry, Principles and Structures</i>. New York: John Willey and Sons.</li> <li>3. Chang, Raymond. 2005. <i>General Chemistry The Essential Concepts Third Edition</i>. USA: McGraw Hill.</li> </ol>
	<b>Additional :</b>	
		<ol style="list-style-type: none"> <li>1. Achmad, Hiskia dan Tupamahu. 1990. <i>Penuntun Belajar Struktur Atom, Struktur Molekul, Sistem Periodik</i>. Bandung: ITB.</li> <li>2. Achmad, Hiskia dan Tupamahu. 1991. <i>Stoikiometri dan Energetika Kimia</i>, Bandung, PT Citra Aditya Bakti.</li> </ol>

	3. Ahmad, Hiskia. 1990. <i>Kimia Larutan</i> . Bandung: Jurusan Kimia FMIPA ITB						
<b>Lecturer</b>	Dr. Harun Nasrudin, M.S.; Dr. Utiya Aizah, M.Pd.; Rusly Hidayah, S.Si., M.Pd.; Prof. Suyatno, M.Si.; Dr. Maria Monica SBW, M.Si.; Dr. Nuniek Herdyastuti, M.Si.; Dr. Amaria, M.Si., Prof. Sari Edy C. M.Si.						
<b>Prerequisite courses</b>	–						
<b>Meeting</b>	<b>The final ability of each activity</b>	<b>Assessment</b>		<b>Learning Forms, Learning Methods, Student Assignment</b>		<b>Reference</b>	<b>Rating Weight (%)</b>
		<b>Indicator</b>	<b>Criteria &amp; Form</b>	<b>Offline</b>	<b>online</b>		
<b>(1)</b>	<b>(2)</b>	<b>(3)</b>	<b>(4)</b>	<b>(5)</b>	<b>(6)</b>	<b>(7)</b>	<b>(8)</b>
<b>1</b>	Describes chemistry as the result of scientific activities that study matter with universal properties	1. Describe scientific steps 2. Describe the extensive and intensive nature 3. Describe the differences in chemical and physical properties, elements, compounds, and mixtures	Essay Writing Test	Interactive discussion	–	<b>Introduction:</b> The stages of the scientific method, Chemistry as a scientific activity, material and energy, extensive and intensive properties, chemical and physical properties, elements, compounds, and mixtures	<b>10</b>
<b>2</b>	Applying the things that underlie stoichiometry, namely: basic laws of chemistry, atoms and molecules, the concept of moles and Avogadro's constant, compound formulas, chemical reactions and polarity and equivalence	1. Describe the basic laws of chemistry 2. Describe the difference between Atom, Molecule, and Molecular Concept	Essay Writing Test	Interactive discussion and exercise	–	<b>Stoichiometry:</b> Basic Chemistry Law, Atoms and Molecules, Mole Concepts, Avogadro Constanta, Compound Formulas,	<b>15</b>



	to complete chemical calculations					Chemical Reactions and Equalization, Polarity and Equivalents	
3	Applying the things that underlie stoichiometry, namely: basic laws of chemistry, atoms and molecules, the concept of moles and Avogadro's constant, compound formulas, chemical reactions and polarity and equivalence to complete chemical calculations	<ol style="list-style-type: none"> <li>1. Applying Avogadro's Constants and Compound Formulas</li> <li>2. Applying Chemical Reactions and Equivalents, Polarities and Equivalents in practice questions</li> </ol>	Essay Writing Test	Interactive discussion and individual task	–	<b>Stoichiometry:</b> Basic Chemistry Law, Atoms and Molecules, Mole Concepts, Avogadro Constanta, Compound Formulas, Chemical Reactions and Equalization, Polarity and Equivalents	
4	Applying the things that underlie stoichiometry, namely: basic laws of chemistry, atoms and molecules, the concept of moles and Avogadro's constant, compound formulas, chemical reactions and polarity and equivalence to complete chemical calculations	<ol style="list-style-type: none"> <li>1. Report how to use and operate equipment according to basic chemistry practicum</li> <li>2. Conduct chemical separation experiments, Laovisier Law and chemical reactions by applying the principles of</li> </ol>	<ul style="list-style-type: none"> <li>▪ Presentation assessment sheet</li> <li>▪ Assessment report laboratory activities</li> </ul>	Presentation, Question and answer, Laboratory activities	–	<b>Stoichiometry:</b> Basic Chemistry Law, Atoms and Molecules, Mole Concepts, Avogadro Constanta, Compound Formulas, Chemical Reactions and Equalization,	

		occupational safety and health				Polarity and Equivalents	
5	Analyzed the development of the discovery and the elementary particles of the atom according to Rutherford, Bohr, wave mechanics and electron configuration	1. Describe the basic particles that make up the atom 2. Analyze the development of atomic theory	<ul style="list-style-type: none"> <li>▪ Essay Writing Test</li> <li>▪ Presentation assessment sheet</li> </ul>	Group task Presentation Question and answer	–	<b>Atomic Structure:</b> Basic Particles, Hydrogen Atom Spectrum and Rutherford Atomic Model, Bohr Atomic Model, Atomic Wave Mechanics Model, Electron Configuration	<b>10</b>
6	Analyzed the development of the discovery and the elementary particles of the atom according to Rutherford, Bohr, wave mechanics and electron configuration	1. Determine the quantum numbers of various atoms 2. Determine the electron configurations of various atoms	Essay Writing Test	Interactive discussion and exercise	–	<b>Atomic Structure:</b> Basic Particles, Hydrogen Atom Spectrum and Rutherford Atomic Model, Bohr Atomic Model, Atomic Wave Mechanics Model, Electron Configuration	
7	Analyze the development, usefulness, and basis for composing the periodic system and its relation to the electron configuration of the elements and their periodic properties	1. Describe the development of the Periodic System of the Elements and electron configuration relationships.	<ul style="list-style-type: none"> <li>▪ Essay Writing Test</li> <li>▪ Presentation assessment sheet</li> </ul>	Group task Presentation Question and answer	–	<b>Periodic System of Elements:</b> Development of the Periodic System, Periodic System and Electron	<b>10</b>

		2. Analyze various characteristics of periodicity				Configuration, Periodicity Properties (Atomic Radius, Ionization Energy, Electron Affinity, and Electronegativity)	
<b>8</b>	<b>Midterm Exams</b>						
<b>9</b>	Identify the relationship between chemical bonds and chemical forces to explain knowledge according to the study program.	<ol style="list-style-type: none"> <li>1. Determine Ionic Bonds, Covalent Bonds, Energy Bonds, and Other Chemical Bonds (van.der Waals, Hydrogen Bonds, Metal Bonds) and their relation to the properties of substances</li> <li>2. Describe the resonance structure of a molecule</li> </ol>	Essay Writing Test	Interactive discussion	–	<b>Chemical Bonds:</b> Ion Bonds, Covalent Bonds, Molecular Structures, Metal Bonds, and Chemical Styles (London Style v.d Waals, Hydrogen Bonds)	<b>15</b>
<b>10</b>	Identify the relationship between chemical bonds and chemical forces to explain knowledge according to the study program.	<ol style="list-style-type: none"> <li>1. Determine the shape and polarity of a molecule based on the Valence Shell Electron Pair Repulsion Theory or hybridization theory.</li> <li>2. Determine the bond order by means of a diagram of the energy levels of the</li> </ol>	Essay Writing Test	Interactive discussion and group task	–	<b>Chemical Bonds:</b> Ion Bonds, Covalent Bonds, Molecular Structures, Metal Bonds, and Chemical Styles (London Style v.d Waals, Hydrogen Bonds)	

		orbitals of various diatomic molecules					
<b>11</b>	Describing terms, the laws of thermodynamics, and determining the occurrence of thermodynamic reactions	<ol style="list-style-type: none"> <li>1. Mendeskripsikan perbedaan Sistem, lingkungan, fungsi keadaan, proses adiabatik, proses isoterm, kerja, dan kapasitas kalor.</li> <li>2. Menerapkan Hukum Termodinamika I, Hukum Hess, dan Energi Ikatan dalam perhitungan</li> <li>3. Describe the differences in systems, environments, state functions, adiabatic processes, isothermic processes, work, and heat capacity.</li> <li>4. Applying the Law of Thermodynamics I, Hess's Law, and Bond Energy in calculations</li> </ol>	Essay Writing Test	Interactive discussion	–	<b>Energetics:</b> Several Terms (Systems, environment, state functions, adiabatic processes, isotherm processes, work, heat capacity, etc.), Law I Thermodynamics, Hess Law, Bonding Energy, Thermochemistry, Law II Thermodynamics, Entropy, Free Energy.	<b>20</b>
<b>12</b>	Describing terms, the laws of thermodynamics, and determining the occurrence of thermodynamic reactions	<ol style="list-style-type: none"> <li>1. Applying Thermochemical equations, Law of Thermodynamics II, Entropy, Free Energy in calculations.</li> </ol>	<ul style="list-style-type: none"> <li>▪ Essay Writing Test</li> <li>▪ Assessment report laboratory activities</li> </ul>	Interactive discussion and Laboratory activities	–	<b>Energetics:</b> Several Terms (Systems, environment, state functions, adiabatic	

		2. Perform thermochemical experiments				processes, isotherm processes, work, heat capacity, etc.), Law I Thermodynamics, Hess Law, Bonding Energy, Thermochemistry, Law II Thermodynamics, Entropy, Free Energy.	
<b>13</b>	Analyze several aspects of the solution and apply them in quantitative terms	<ol style="list-style-type: none"> <li>1. Calculating the various concentrations of the solution</li> <li>2. Determine the colligative properties of electrolyte and non-electrolyte solutions.</li> <li>3. Distinguishing the acid-base theory</li> </ol>	Essay Writing Test	Interactive discussion and group task	–	<b>Solution:</b> Electrolyte and non-electrolyte solution, colligative properties, acid-base, pH of solution, hydrolysis, namesake ion, buffer solution, and titration.	<b>20</b>
<b>14</b>	Analyze several aspects of the solution and apply them in quantitative terms	<ol style="list-style-type: none"> <li>1. Calculate the pH of the solution.</li> <li>2. Analyze the ionic equilibrium in the salt solution and relate the pH.</li> <li>3. Determine the working principle, pH calculation and the</li> </ol>	Essay Writing Test	Interactive discussion and group task	–	<b>Solution:</b> Electrolyte and non-electrolyte solution, colligative properties, acid-base, pH of solution, hydrolysis,	

		role of buffer solutions in life.				namesake ion, buffer solution, and titration.	
15	Analyze several aspects of the solution and apply them in quantitative terms	<ol style="list-style-type: none"> <li>1. Determine the pH indicator route.</li> <li>2. Analyze data on the results of various types of acid-base titrations</li> <li>3. Carry out an acid-base titration experiment</li> </ol>	<ul style="list-style-type: none"> <li>▪ Essay Writing Test</li> <li>▪ Assessment report laboratory activities</li> </ul>	Interactive discussion and Laboratory activities	–	<b>Solution:</b> Electrolyte and non-electrolyte solution, colligative properties, acid-base, pH of solution, hydrolysis, namesake ion, buffer solution, and titration.	
16	<b>Final Exams</b>						<b>100</b>

#### A.4. MAPPING OF LEARNING OUTCOMES – COURSE OUTCOMES

##### A.4.1. The Expected Program Learning Outcomes (PLO) of Undergraduate Program of Education Chemistry (UPCE)

NO	ASPECTS	PLO	CODE
1	KNOWLEDGE	1. Capable to demonstrate knowledge related to theoretical concepts about structure, dynamics, and energy, as well as the basic principles of separation, analysis, synthesis and characterization of chemicals	KNO-1
		2. Capable to demonstrate the pedagogical knowledge of chemistry in designing, implementing, and evaluating chemistry learning	KNO-2
2	SKILL	3. Mastering the principles of occupational health and safety, managing laboratories, using the equipment and operating chemical instruments	SKI-1
		4. Capable to design, implement, evaluate, learn and develop chemistry learning media by utilizing Information and Communication Technology	SKI-2
3	COMPETENCIES	5. Applying logical, critical, systematic and innovative thinking in the context of development or implementation of science, technology, and art that regards and applies humanities in accordance with chemistry education in solving problems	COM-1
		6. Mastering the basics of the scientific method, designing and conducting research, writing scientific reports and communicating them both verbally and in writing by utilizing information and communication technology in the field of education	COM-2
4	ATTITUDE AND SOCIAL	7. Capable to make decisions based on data/information in order to complete their responsibility assignment and evaluate the performance that has been done both individually and in groups, have an entrepreneurial spirit with environmental insight	SOC-1
		8. Capable to adapt to various developments in chemistry, develop and learn continuously throughout life to continue education, both formal and informal	SOC-2

#### A4.2. The Education Program Objectives (PEOs) of Basic Chemistry I.

PEO 1. Comprehending the concept and chemistry learning, laboratory management, scientific method, and ICT as well as its implementation to solve the problem in their profession.

PEO 5. Having capability to develop and apply chemistry concept along with the progress of science and technology as well as humanities values.

#### A4.3. Mapping of Program Learning Outcomes (PLO) – Education Program Objectives (PEOs)

	PLO 1 (KNO-1)	PLO 3 (SKI-1)
PEO 1		
PEO 5		

### B. COURSE ASSESSMENT

#### B.1. Assessment Rubric

Cognitive Criteria

1. The ability to give answers correctly
2. The ability to provide argumentation according to theory
3. The ability to provide systematic explanations
4. The ability to solve problems comprehensively

#### B.2. Assessment System

Final Assessment Course with practicum

Practicum	: 20%
Group/Individuals Assignment	: 20%
Midterm examination	: 30%
Final examination	: 30%

#### Distribution of the weight of the ability of the test item

	PLO 1 (KNO-1)	PLO 3 (SKI-1)	Total
Practicum	60%	40%	100%
Group/Individuals Assignment	70%	30%	100%
Midterm examination	70%	30%	100%
Final examination	80%	20%	100%

#### Success Criteria of Program Learning Outcomes (PLO)

Excellent	80
Good	70
Satisfy	55
Failed	< 55

Final index for undergraduate program defined as follow:



Final Index	Range
A	4 (85 - 100)
A <sup>-</sup>	3,75 (80 - 85)
B <sup>+</sup>	3,5 (75 - 80)
B	3 (70 - 75)
B <sup>-</sup>	2,75 (65 - 70)
C <sup>+</sup>	2,5 (60 - 65)
C	2 (55 - 60)
D	1 (40 - 55)
E	0 (0 - 40)

## C. COURSE DEVELOPMENT

### C.1. Academic Year 2019/2020 odd semester

Parameter	of person	Percentage
Number of students taking this subject	87	100 %
Number of students who pass at first attempt (>B <sup>-</sup> )	68	78,16 %
Number of students who pass at first attempt (C - B <sup>-</sup> )	18	20,69 %
Number of failed students after remedial (D & E)	1	1,15 %

### C.2. Problems Analysis

In 2019/2020 academic year in the basic chemistry I course, there were 100 % of students who had passed the examination at the first attempt. At the end of the semester examination, there is no remedial. There is one student who did not graduate because the student did not take the final exam and collecting assignments. There are 18 students who graduated, but the grades are below standard, namely 55 - 70. So, it was thought that the learning strategy/methods still need to be improved to achieve higher results in the future. The average final score in 2019/2020 is lower than before, due students have different characteristics, namely they difficult to cooperate with their group and not serious when doing the task, therefore the have lack of average score.

### C.3. Solutive Strategy

New teaching and learning methods should be developed for the next academic years, consisting of:

1. Redesigning the course material (PPT slides, course contents, etc.) to become more interesting and interactive to stimulate student's interest in this course.
2. Giving "lecture by online" to stimulate our students to learn about the next lecture topics.
3. Enhance the cooperative skills of students with exchange the methods and models of learning

## D. APPENDICES

### D.1. DOCUMENT OF COURSE ACTIVITY

#### D.1.1. Lecture's journal and student's attendance form [siakadu.uneca.ac.id](http://siakadu.uneca.ac.id)

Siakadu | Universitas ...

Jurnal Pembelajaran Kelas Dosen | KELAS 2019C - 01 Pendidikan Kimia

Jika Peserta Perkuliahan masih 0 maka presensi belum di klik simpan, segera simpan presensi pada pertemuan tersebut.

Sahibernal Jani Kusni

Pertemuan	Tanggal / Kelas	Topik	Status	Peserta	TGT/SEMPSU	Peserta	Barcode
04.1	27 Agustus 2019 Dosen: HARIBU NASRUDDIN	Menentukan tetapan termodinamika dengan menggunakan persamaan keadaan untuk gas nyata dan menggunakan persamaan keadaan untuk gas nyata dan persamaan keadaan lainnya.	Denda	01			
04.2	28 Agustus 2019 Dosen: LITVA ARIYATI	20. Menentukan hukum kekekalan energi dalam sistem termodinamika dan menggunakan persamaan keadaan untuk gas nyata dan persamaan keadaan lainnya.	Terjadwal	02			
04.3	29 Agustus 2019 Dosen: HARIBU NASRUDDIN	20. Menentukan hukum kekekalan energi dalam sistem termodinamika dan menggunakan persamaan keadaan untuk gas nyata dan persamaan keadaan lainnya.	Terjadwal	03			
04.4	30 Agustus 2019 Dosen: HARIBU NASRUDDIN	20. Menentukan hukum kekekalan energi dalam sistem termodinamika dan menggunakan persamaan keadaan untuk gas nyata dan persamaan keadaan lainnya.	Terjadwal	04			
04.5	31 Agustus 2019 Dosen: HARIBU NASRUDDIN	20. Menentukan hukum kekekalan energi dalam sistem termodinamika dan menggunakan persamaan keadaan untuk gas nyata dan persamaan keadaan lainnya.	Terjadwal	05			
04.6	01 September 2019 Dosen: HARIBU NASRUDDIN	20. Menentukan hukum kekekalan energi dalam sistem termodinamika dan menggunakan persamaan keadaan untuk gas nyata dan persamaan keadaan lainnya.	Terjadwal	06			
04.7	02 September 2019 Dosen: LITVA ARIYATI	20. Menentukan hukum kekekalan energi dalam sistem termodinamika dan menggunakan persamaan keadaan untuk gas nyata dan persamaan keadaan lainnya.	Terjadwal	07			
04.8	03 September 2019 Dosen: HARIBU NASRUDDIN	20. Menentukan hukum kekekalan energi dalam sistem termodinamika dan menggunakan persamaan keadaan untuk gas nyata dan persamaan keadaan lainnya.	Terjadwal	08			
04.9	04 September 2019 Dosen: LITVA ARIYATI	20. Menentukan hukum kekekalan energi dalam sistem termodinamika dan menggunakan persamaan keadaan untuk gas nyata dan persamaan keadaan lainnya.	Terjadwal	09			
04.10	05 September 2019 Dosen: LITVA ARIYATI	20. Menentukan hukum kekekalan energi dalam sistem termodinamika dan menggunakan persamaan keadaan untuk gas nyata dan persamaan keadaan lainnya.	Terjadwal	10			
04.11	06 September 2019 Dosen: HARIBU NASRUDDIN	20. Menentukan hukum kekekalan energi dalam sistem termodinamika dan menggunakan persamaan keadaan untuk gas nyata dan persamaan keadaan lainnya.	Terjadwal	11			
04.12	07 September 2019 Dosen: HARIBU NASRUDDIN	20. Menentukan hukum kekekalan energi dalam sistem termodinamika dan menggunakan persamaan keadaan untuk gas nyata dan persamaan keadaan lainnya.	Terjadwal	12			
04.13	08 September 2019 Dosen: LITVA ARIYATI	20. Menentukan hukum kekekalan energi dalam sistem termodinamika dan menggunakan persamaan keadaan untuk gas nyata dan persamaan keadaan lainnya.	Terjadwal	13			
04.14	09 September 2019 Dosen: LITVA ARIYATI	20. Menentukan hukum kekekalan energi dalam sistem termodinamika dan menggunakan persamaan keadaan untuk gas nyata dan persamaan keadaan lainnya.	Terjadwal	14			
04.15	10 September 2019 Dosen: LITVA ARIYATI	20. Menentukan hukum kekekalan energi dalam sistem termodinamika dan menggunakan persamaan keadaan untuk gas nyata dan persamaan keadaan lainnya.	Terjadwal	15			



KEMENTERIAN PENDIDIKAN DAN KEBUDAYAAN  
UNIVERSITAS NEGERI SURABAYA

Jl. Lidah Wetan, Surabaya - 60213  
Telepon : +6231-95121932  
Faksimile : +6231-95424932  
e-mail : bskpk@unesa.ac.id

**PRESENSI KULIAH**  
Periode 2019/2020 Gasal

Mata Kuliah : Kimia Dasar I  
Kelas : 2019C  
Prodi : S1 Pendidikan Kimia

Dosen : Dr. Harun Nasrudin, M.S.  
Dr. Utilya Azizah, M.Pd.

No	NIM	Nama Mahasiswa	Pertemuan Ke															%
			1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	
			23 Aug 19	26 Aug 19	02 Sep 19	09 Sep 19	16 Sep 19	23 Sep 19	30 Sep 19	07 Oct 19	14 Oct 19	21 Oct 19	28 Oct 19	04 Nov 19	11 Nov 19	18 Nov 19	25 Nov 19	
1.	19030194001	EKA NUR ARIYANTI	H	H	H	H	H	H	H	H	H	H	H	H	H	H	100 %	
2.	19030194005	DINI ANGGRAINI	H	H	H	H	H	H	H	H	H	H	H	H	H	H	100 %	
3.	19030194007	SELVIA NUR'AINI	H	H	H	H	H	H	H	H	H	H	H	H	H	H	100 %	
4.	19030194009	AMALIA CAHYANING WULAN ACUSTINE	H	H	H	H	H	H	H	H	H	H	H	H	H	H	100 %	
5.	19030194015	TITO VANZAL	H	H	H	H	H	H	H	H	H	H	H	H	H	H	100 %	
6.	19030194016	DWI WILLUENG FATTIKASARI	H	H	H	H	H	H	H	H	H	H	H	H	H	H	100 %	
7.	19030194018	AZZA NURIAH WIDOWATI	H	H	H	H	H	H	H	H	H	H	H	H	H	H	100 %	
8.	19030194022	HANY ARMAYANTI	H	H	H	H	H	H	H	H	H	H	H	H	H	H	100 %	
9.	19030194023	ILMIATUL MUFALDAH	H	H	H	H	H	H	H	H	H	H	H	H	H	H	100 %	
10.	19030194025	DIAN ZULFATUR RIZOYAH	H	H	H	H	H	H	H	H	H	H	H	H	H	H	100 %	
11.	19030194028	ADELIA HOURISTA KHAININIZA	H	H	H	H	H	H	H	H	H	H	H	H	H	H	100 %	
12.	19030194029	NOVITA INDAH RAMADHANI	H	H	H	H	H	H	H	H	H	H	H	H	H	H	100 %	
13.	19030194032	SEPTIA NURKHALIDA	H	H	H	H	H	H	H	H	H	H	H	H	H	H	100 %	
14.	19030194034	AFIQA AZRA AMANINA	H	H	H	H	H	H	H	H	H	H	H	H	H	H	100 %	
15.	19030194037	SALSABILA ALMAS DWI RAN II	H	H	H	H	H	H	H	H	H	H	H	H	H	H	100 %	
16.	19030194041	META KHIL JANAH	H	H	H	H	H	H	H	H	H	H	H	H	H	H	100 %	
17.	19030194043	NIRMALA PUTERI PATARI	H	H	H	H	H	H	H	H	H	H	H	H	H	H	100 %	
18.	19030194045	MUHAMMAD DANU ERLANGGA	H	H	H	H	H	H	H	H	H	H	H	H	H	H	100 %	
19.	19030194047	BELLA WAJIPUNING TYAS	H	H	H	H	H	H	H	H	H	H	H	H	H	H	100 %	
20.	19030194048	FAUZIA HANIM ZULFAH	H	H	H	H	H	H	H	H	H	H	H	H	H	H	100 %	
21.	19030194050	EKA HASLUNDA FATMAWATI	H	H	H	H	H	H	H	H	H	H	H	H	H	H	100 %	
22.	19030194055	ELFA SELVIANA	H	H	H	H	H	H	H	H	H	H	H	H	H	H	100 %	
23.	19030194059	SUDZULASMAIS	H	H	H	H	H	H	H	H	H	H	H	H	H	H	100 %	
24.	19030194060	AINUN TAZKIA	H	H	H	H	H	H	H	H	H	H	H	H	H	H	100 %	
25.	19030194058	SABRINA AJI SABILA	H	H	H	H	H	H	H	H	H	H	H	H	H	H	100 %	
26.	19030194056	RYO WIDI DANIELSON	H	H	H	H	H	H	H	H	H	H	H	H	H	H	100 %	
27.	19030194076	SISGA W/ DIANA PLTRI	H	H	H	H	H	H	H	H	H	H	H	H	H	H	100 %	
28.	19030194077	AIZA ALYA	H	H	H	H	H	H	H	H	H	H	H	H	H	H	100 %	
29.	19030194085	RINTIS MEGA AYIRAHMA	H	H	H	H	H	H	H	H	H	H	H	H	H	H	100 %	
Tanda Tangan Dosen / Asisten																		

**D.1.2. Sample of statement of examination official report**

(Scan Berita Acara Ujian Kimia Dasar 1)

## D.2. SAMPLE OF STUDENT WORK

### D.2.1. Sample of Test Paper



KEMENTERIAN RISET, TEKNOLOGI DAN PENDIDIKAN TINGGI  
UNIVERSITAS NEGERI SURABAYA  
FAKULTAS MATEMATIKA DAN ILMU PENGETAHUAN ALAM  
JURUSAN KIMIA

Kampus Ketintang  
Jalan Ketintang Gedung C5 dan C6  
Surabaya 60231  
T: +6231-8298761  
F: +6231-8298761



Management System  
ISO 9001:2015  
www.tuv.com  
ID: 8108860021

#### FINAL TEST OF ODD SEMESTER 2019/2020

Examination Subject	: Basic Chemistry I
Department/Faculty	: Chemistry/Mathematics and Natural Sciences
Program/Year	: Chemistry Education (PKU 2019)
Day / date	: Wednesday/18 December 2019
Period	: 100 minutes
Time	: III
Lecturers	: Team
Characteristic	: Closed Book

#### Directions:

1. Answer the following questions on the answer sheet.
2. Used a calculator (not a handphone)

#### A. CHEMICAL BONDING (score 40)

1. Consider the physical properties of the following two substances

Num	Physical Properties	Substance A	Substance B
1	The electric conductivity melted	Conduct electricity	Does not Conduct electricity
2	The electrical conductivity of solutions	Conduct electricity	Does not Conduct electricity
3	Boiling point and melting point	high	low

Based on these data, determine the types of bonds contained in substance A and substance B.

2. There are two elements with the notation 12A and 35B. If the two elements are bonded, determine the shape of the molecule and the polarity that occurs.
3. Using the molecular orbital theory, determine the bond order  $O_2^{2-}$  if the atomic number is  $O = 8$  by drawing the energy level diagram.

#### B. ENERGETICS (Score 20)

4. Find  $H$  in the reaction  $H_2C = C_2(g) + H_2(g) \rightarrow H_3C - C_2(g)$ , if the bond energy is  $C = C = 615$  kJ/mol;  $C - H = 414,2$  kJ/mol;  $H - H = 436$  kJ/mol;  $C - C = 347,3$  kJ/mol.
5. Consider the methane combustion reaction:  $CH_4(g) + O_2(g) \rightarrow CO_2(g) + 2H_2O(l)$   
If you know the price of the change in entropy is  $-242,2$  J/K mole and the change in enthalpy  $-890,4$  kJ/mole, calculate the standard Gibbs free energy change at 250C? The reaction takes place spontaneously or not?

#### C. SOLUTION (Score 40)

6. Determine the mass (in grams) of glucose (molar mass = 180.2 g / mol) needed to make 1.00 liters of intravenous injection solution whose osmotic pressure is the same as the osmotic pressure of blood, if the osmotic pressure of the blood is 7.65 atm at 37°C.
7. Calculate the pH of the solution made by mixing 100 mL NaOH 0,200 M and 150 mL  $CH_3COOH$  0,400 M ( $K_a = 1,8 \times 10^{-5}$ ).
8. 1.7872 grams of the sample contains sodium carbonate. The sample is dissolved up to 100 mL. 25 mL of the sample is titrated with indicator A and requires 21.35 mL of 0.1 M HCl until the indicator changes color. (Known  $K_a H_2CO_3 = 4,2 \cdot 10^{-7}$ ;  $K_a NaHCO_3 = 4,8 \cdot 10^{-11}$ ). Specify: a) pH at the equivalence point, and b) The indicator used in the titration.

Note: Ar C = 12,01 ; H = 1 ; Na = 22,99; Cl = 35,45 ; O = 15,99

D.2.2. Sample of Student's Work

Nama	: ELFA SELVIANA	Tanda Tangan
Jurusan	: KIMIA (PKC 2019)	
No. Reg.	: 19030194059	
Mata Kuliah	: KIMIA DASAR	Nilai
Dosen	: IBU UTIYA AZIZAH	89
Hari/Tanggal	: Rabu/18 Dec. 2019	

A. IKATAN KIMIA

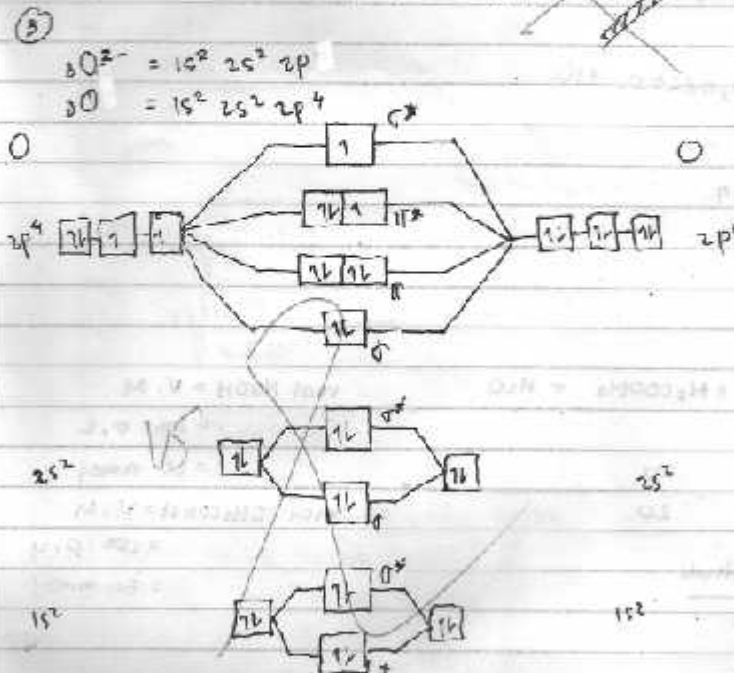
1. ikatan pada zat A = 10ri  
ikatan pada zat B = kovalen nonpolar

2. 15A → 1 2, 8, 2 → ingin stabil ≈ 10Ne A<sup>2+</sup>  
35B → 2, 8, 16, 8, 1 → ingin stabil ≈ 36Kr B<sup>-</sup>

8 A<sup>2+</sup> + B<sup>-</sup> → AB<sub>2</sub> → ikatan ~~ion~~

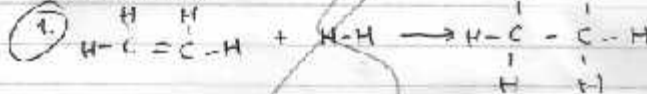
$$PEB = \frac{2-2}{2} = 0$$

AX<sub>2</sub> = linier



$$\text{Orde ikatan} = \frac{10 - 8}{2} = \frac{2}{2} = 1$$

B. ENERGETIKA



$\Delta H = \text{Kiri} - \text{Kanan}$

43  
35  
11  
89



②.  $\Delta S = -242,2 \text{ J/Kmol}$   
 $\Delta H = -890,4 \text{ kJ/mol} = -890,4 \cdot 10^3 \text{ J/mol}$   
 $T = 25^\circ\text{C} = 298 \text{ K}$

Ditanya:  $\Delta G$

$$\Delta G = \Delta H - (T \cdot \Delta S)$$

$$= -890,4 \cdot 10^3 - (298 \cdot (-242,2))$$

$$= -890,4 \cdot 10^3 + 72.175,6$$

$$= -818,224,4 \text{ J/mol} = -818,2 \text{ kJ/mol}$$

reaksi spontan

C. LARUTAN

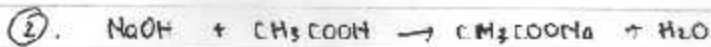
①.  $\pi = M \cdot R \cdot T$

$$\pi = \frac{\text{gram}}{M_r} \cdot \frac{1000}{P} \cdot R \cdot T$$

$$7,65 = \frac{\text{gram}}{180,2} \cdot \frac{1000}{1000} \cdot 0,08205 \cdot 310$$

$$7,65 = \frac{\text{gram}}{180,2} \cdot 25,44$$

$$54,2 = \text{gram}$$



m	20	60	
r	20	20	20
s	-	40	20

mol NaOH = V · M  
 $= 100 \cdot 0,2$

$= 20 \text{ mmol}$

mol CH<sub>3</sub>COOH = V · M

$= 150 \cdot 0,4$

$= 60 \text{ mmol}$

$[\text{H}^+] = K_a \cdot \frac{n \text{ asam lemah}}{n \text{ garam}}$

$$= 1,8 \cdot 10^{-5} \cdot \frac{40}{20}$$


$$= 3,6 \cdot 10^{-5}$$

pH =  $5 - \log 3,6$

$= 5 - 0,556$

$= 4,443$

## UTS/UAS FAKULTAS MIPA - UNESA

Nama : ELFA SEWIANA	Tanda Tangan
Jurusan : KIMIA (PKC 2019)	
No. Reg. : 19030194055	
Mata Kuliah : KIMIA DASAR	Nilai
Dosen : IBU UTIYA ALIZAH	
Hari/Tanggal : Rabu, 18 Des. 2019	

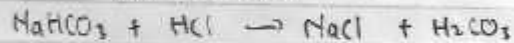
3. a. PH pada titik ekuivalen

$$M_1 \cdot V_1 = M_2 \cdot V_2$$

$$25 \cdot M_1 = 0,1 \cdot 21,35$$

$$M_1 = \frac{0,1 \cdot 21,35}{25}$$

$$M_1 = 0,0854$$



$$m \quad 2,135 \quad 2,135$$

$$F \quad 2,135 \quad 2,135 \quad 2,135$$

$$S \quad - \quad - \quad 2,135 \text{ mmol}$$

$$\cdot \text{ mol NaHCO}_3 = M \cdot V$$

$$= 0,0854 \cdot 25$$

$$= 2,135 \text{ mmol}$$

$$\cdot \text{ mol HCl} = 21,35 \cdot 0,1$$

$$= 2,135 \text{ mmol}$$

$$\rightarrow M = \frac{2,135}{21,35 + 25} = \frac{2,135}{46,35} = 0,046$$

$$[\text{OH}^-] = \sqrt{\frac{K_w}{K_a} \cdot [G]}$$

$$= \sqrt{\frac{10^{-14}}{4,8 \cdot 10^{-11}} \cdot 0,046}$$

$$= \sqrt{9,583 \cdot 10^{-6}}$$

$$= 3,096 \cdot 10^{-3}$$

$$\text{pOH} = 3 - \log 3,096$$

$$\text{PH} = 14 - (3 - \log 3,096)$$

$$= 11 + \log 3,096$$

$$= 11 + 0,49$$

$$= 11,49$$



## D.3. RECAPITULATION OF ASSESSMENT

### D.3.1. Validate Test Item

The end-of-semester evaluation questions consist of eight items in the form of essay questions analyzed content through experts in the appropriate field of Chemistry Education analyzed. Essay questions are validated with expert judgment in the course team members. The analysis was conducted by taking into account several aspects, namely the suitability of the questions with the course outcome, language, content and construct.

### D.3.2 Evaluation Results of Basic Chemistry I

PROGRAM STUDI SL Pendidikan Kimia  
 DAFTAR NILAI MAHASISWA  
 Mata Kuliah : Kimia Dasar I  
 Kelas : 2019C  
 Tahun Ajaran : 2019/2020 Gasal

Original data :



**Keterangan :**

1. Kelompok nilai yang nilai hanya : Part, Tugas, UTS dan UAS
2. Nilai UAS mahasiswa dengan kehadiran dibawah 73.3% (kolom dg warna merah) tidak akan diinput
3. Jangan merubah apapun di dokumen ini kecuali pada point nomor satu di atas.
4. PPTI / BAAK tidak menerima file nilai untuk diupload. Proses upload nilai dilakukan oleh dosen pengampu yang bersangkutan.

No	NIM	Nama Mahasiswa	Angkatan	Kehadiran	Prakt	Tugas	UTS	UAS	NA	Huruf	Pakai
1	19030194003	NANDA FAUZIYAH FEBRIANTI	2019	100%	83	75	45	62	66.7	B-	1
2	19030194004	VERAWATI ISNAINI	2019	100%	86	76	57	51	66.7	B-	1
3	19030194008	ANNISA PUTRIA DEWITASARI	2019	100%	94	86	60	75	79.1	B+	1
4	19030194010	ANDANG NURHUDA	2019	100%	91	83	55	68	74.5	B	1
5	19030194011	HASNA CHOIRIYAH	2019	100%	70	62	45	75	64.1	C+	1
6	19030194012	ALIMATHUS SA'DIYAH	2019	100%	82	74	55	74	71.8	B	1
7	19030194013	RIA FATMAWATI	2019	100%	73	65	50	62	62.7	C+	1
8	19030194021	EVITA HARTI NANDA	2019	100%	90	82	60	83	79.5	B+	1
9	19030194024	ANNISA NUR RAHMAWATI	2019	100%	87	79	55	77	75.2	B+	1
10	19030194026	DYAH KIRANI NOVIYANA	2019	100%	80	72	45	80	70.6	B	1
11	19030194030	NAWANG WAHYU WULANDARI	2019	100%	78	68	60	63	66.9	B-	1
12	19030194031	AMALIA MUNJIATUL UMMAH	2019	100%	83	75	55	75	72.6	B	1
13	19030194033	SALSABILA RACHMASARI PUTRI	2019	100%	88	80	65	83	79.5	B+	1
14	19030194036	MAYA KUMALASARI	2019	100%	95	87	70	80	83.1	A-	1
15	19030194039	ARZA RIZKY SEPTI ANSYACH	2019	100%	92	84	45	74	74.8	B	1
16	19030194042	PUTRI NUR AZIZAH	2019	100%	94	86	63	87	83.3	A-	1
17	19030194044	YULISA DWI ANGGRAENI	2019	100%	92	84	45	84	77.8	B+	1
18	19030194054	CICI APRILIA	2019	100%	88	80	55	65	72.1	B	1
19	19030194056	FADHILATUR ROCHMATIN	2019	100%	87	79	60	81	77.4	B+	1
20	19030194062	DINI CATUR ANISAH	2019	100%	81	73	67	91	78.8	B+	1
21	19030194064	FITRIA RAHMATUL ULA	2019	100%	84	76	57	87	77.1	B+	1
22	19030194072	ISTI INDRA WINARSEH	2019	100%	91	83	65	80	80.1	A-	1
23	19030194073	SHINTA TAQIYYAH NABILAH NUHA	2019	100%	85	77	45	63	68	B-	1
24	19030194074	SANIYYATUL AWALIYAH	2019	100%	83	75	45	64	67.3	B-	1
25	19030194075	AMELIA WULANDARI	2019	100%	72	64	45	64	61.8	C+	1
26	19030194080	OCTAVIA DWI FADLIILAH	2019	100%	77	69	57	51	62.8	C+	1
27	19030194082	TASSHA PUTRI RATNASARI	2019	100%	84	76	45	74	70.8	B	1
28	19030194083	MAHARANI AGUSTINA ARIVI	2019	100%	83	75	75	67	74.2	B	1
29	19030194084	IZZATUL MUHIDAH	2019	100%	84	76	75	78	78	B+	1
30	19030194087	AGUNG WIJAYA	2019	100%	76	68	65	53	64.5	C+	1
31	19030194002	SINTIA NUR AENI	2019	93.33%	87	82.3	78	80	81.69	A-	1
32	19030194006	MAULIDIA USWATUN KHASANAH	2019	93.33%	89	85.7	70	85	83.01	A-	1
33	19030194014	DWI MEI SILVIA	2019	93.33%	86	81.3	80	73	79.49	B+	1
34	19030194017	SAFIRA FIRDAUS YAHYA	2019	93.33%	88	82	70	87	82.3	A-	1
35	19030194019	KHUROTA A'YUNIN	2019	93.33%	86	84.7	95	81	85.91	A	1

36	19030194020	NENI ANUGRAHENI NURRAHMAH	2019	93.33%	72	80	92	83	81.7	A-	1
37	19030194035	DANANG PUTRA PRATAMA	2019	93.33%	88	57	69	78	71.9	B	1
38	19030194038	SYIFA AMANDHA	2019	93.33%	80	82.7	74	69	76.31	B+	1
39	19030194040	WELLA YEKTI INKOMARA	2019	93.33%	67	76.7	70	58	67.81	B-	1
40	19030194046	ELVIRA MIFTARIDA AFANDI	2019	93.33%	65	76	100	70	76.8	B+	1
41	19030194049	ANNISA NABILA	2019	93.33%	70	77.7	75	68	72.71	B	1
42	19030194051	NADIA EKA VANIA SUNARTO	2019	93.33%	67	76.7	68	71	71.31	B	1
43	19030194052	FAJAR NOVA PRASETYO	2019	93.33%	65	76	68	78	72.8	B	1
44	19030194053	GITA THERESA ARY SUDARSONO	2019	93.33%	68	77	71	75	73.4	B	1
45	19030194057	ZULIA TRIS FEBRIANTI	2019	93.33%	70	76.7	80	78	76.41	B+	1
46	19030194061	NUR LAILIL APRILIA	2019	93.33%	74	80.7	69	68	73.21	B	1
47	19030194063	ALVIN MAGHFIRAH	2019	93.33%	70	79.3	70	79	75.49	B+	1
48	19030194066	COLLIA NAWANG PUTRI	2019	93.33%	66	76.3	80	83	76.99	B+	1
49	19030194067	IGA PUTRI SUBANDI	2019	93.33%	71	77	75	55	68.8	B-	1
50	19030194070	RANI RATNA KUSUMA	2019	93.33%	67	75	82	49	67	B-	1
51	19030194071	FAIZ RIZKY NUR AWWALUDIN	2019	93.33%	65	77.7	72	15	55.21	C	1
52	19030194078	SYARIFAH AISAH	2019	93.33%	68	77	69	62	69.1	B-	1
53	19030194079	YESIKA DWI PRASTIWI	2019	93.33%	73	71.3	72	49	65.09	B-	1
54	19030194081	LILLA PANGESTU HARWYANDANI	2019	93.33%	67	78.3	75	69	72.59	B	1
55	19030194086	ADELLA ICHA ARDHANI	2019	93.33%	73	78.7	80	82	78.81	B+	1
56	19030194088	FIRDA NURIN NIKMAH	2019	93.33%	68	49.3	43	0	36.99	E	1
57	19030194089	FITANI WARDHA MACHFIRO	2019	93.33%	73	75.7	51	49	62.21	C+	1
58	19030194090	TSABITA LATHUF ZHAFIRAH.A	2019	93.33%	70	54.7	68	43	56.91	C	1
59	19030194001	EKA NUR AFIYANTI	2019	100%	77	82	85	75	79.5	B+	1
60	19030194005	DINI ANGGRAINI	2019	100%	84	86	83	89	85.9	A	1
61	19030194007	SELVIA NURAINI	2019	100%	82	84	83	74	80.4	A-	1
62	19030194009	AMALIA CAHYANING WULAN AGUSTINE	2019	100%	78	80	79	77	78.5	B+	1
63	19030194015	TITO VANZAL	2019	100%	89	91	88	89	89.4	A	1
64	19030194016	DWI WILUJENG FATTIKASARI	2019	100%	90	91	78	95	89.4	A	1
65	19030194018	AZZA NURIAH WIDOWATI	2019	100%	78	80	70	64	72.8	B	1
66	19030194022	HANY ARMAYANTI	2019	100%	86	88	83	89	86.9	A	1
67	19030194023	ILMIATUL MUFA'IDAH	2019	100%	87	85	87	91	87.6	A	1
68	19030194025	DIAN ZULFATUR RIZQIYAH	2019	100%	77	79	72	71	74.8	B	1
69	19030194028	ADELIA FOURISTA KHAIRINIZA	2019	100%	75	77	80	70	75.1	B+	1
70	19030194029	NOVITA INDAH RAMADHANI	2019	100%	75	80	79	58	72.2	B	1
71	19030194032	SEPTIA NURKHALIDA	2019	100%	87	89	74	87	85	A	1
72	19030194034	AFIQA AZRA AMANINA	2019	100%	84	86	84	73	81.3	A-	1
73	19030194037	SALSABILA ALMAS DWI RANTI	2019	100%	83	85	73	59	74.4	B	1
74	19030194041	MIFTAKHUL JANAH	2019	100%	77	79	85	41	68.4	B-	1
75	19030194043	NIRMALA PUTERI BATARI	2019	100%	81	83	79	69	77.6	B+	1
76	19030194045	MUHAMMAD DANU ERLANGGA	2019	100%	91	93	80	95	90.6	A	1
77	19030194047	BELLA WAHYUNING TYAS	2019	100%	89	91	81	88	87.7	A	1
78	19030194048	FAUZIA HANIM ZULFAH	2019	100%	88	90	75	83	84.5	A-	1
79	19030194050	EKA HASLINDA FATMAWATI	2019	100%	84	86	77	60	76	B+	1
80	19030194055	ELFA SELVIANA	2019	100%	87	89	82	95	89	A	1
81	19030194059	SUDZUASMAIS	2019	100%	83	85	70	52	71.7	B	1
82	19030194060	AINUN TAZKIA	2019	100%	83	85	76	83	82.2	A-	1
83	19030194068	SABRINA AJI SABILA	2019	100%	85	87	73	82	82.3	A-	1
84	19030194069	RYO WIDI DANIELSON	2019	100%	85	87	80	80	83.1	A-	1
85	19030194076	SISKA WIDIANA PUTRI	2019	100%	80	80	80	89	82.7	A-	1
86	19030194077	AIZA ALYA	2019	100%	82	84	81	85	83.3	A-	1
87	19030194085	RINTIS MEGA AYIRAHMA	2019	100%	65	70	81	65	69.7	B-	1

### D.3.3 Percentage of PLO achievements of basic chemistry I at Academic Year 2019/2020

	PLO-1	PLO-2	PLO-3	PLO-4	PLO-5	PLO-6	PLO-7	PLO-8
EXELENCE	26%		30%					
GOOD	45%		47%					
SATISFY	26%		22%					
FALSE	2%		1%					
	100%	0%	100%	0%	0%	0%	0%	0%

